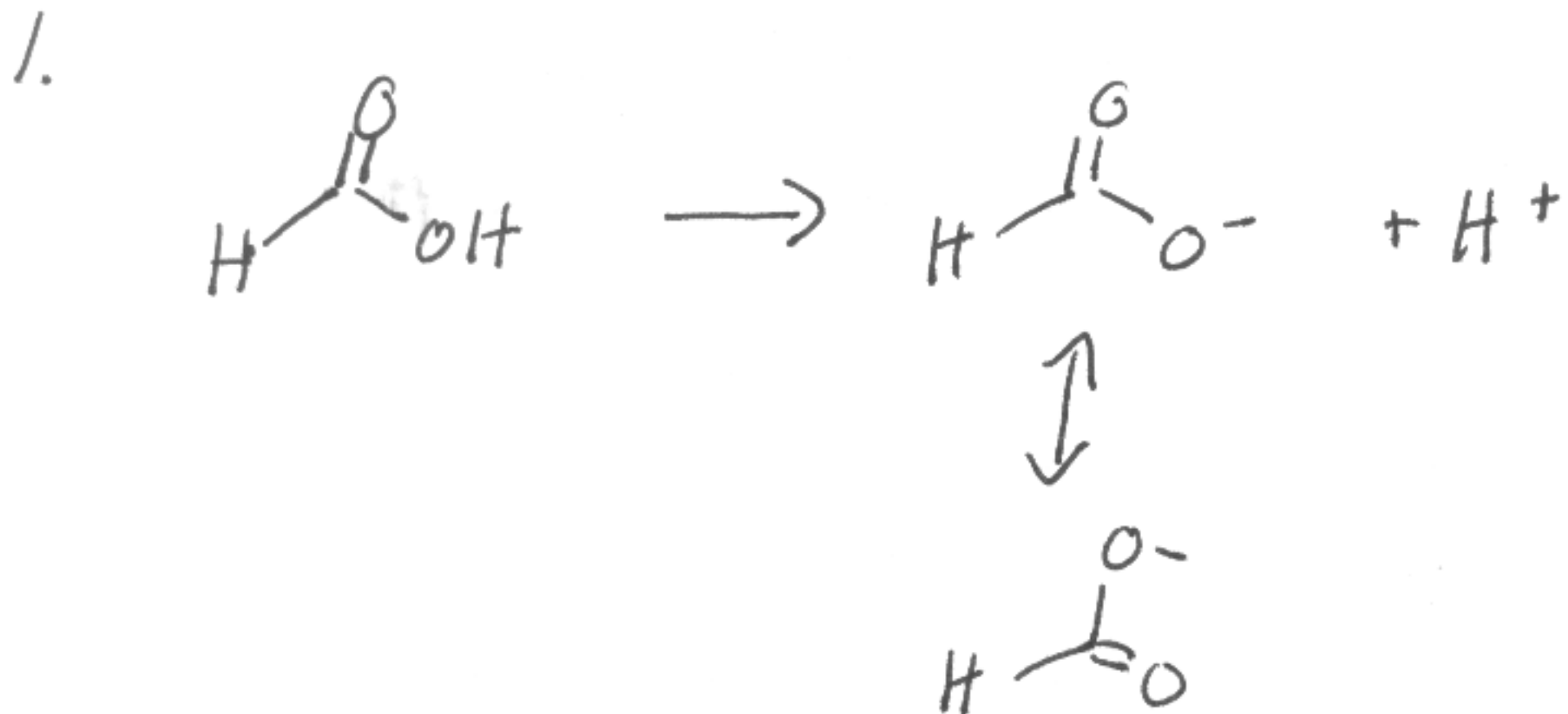



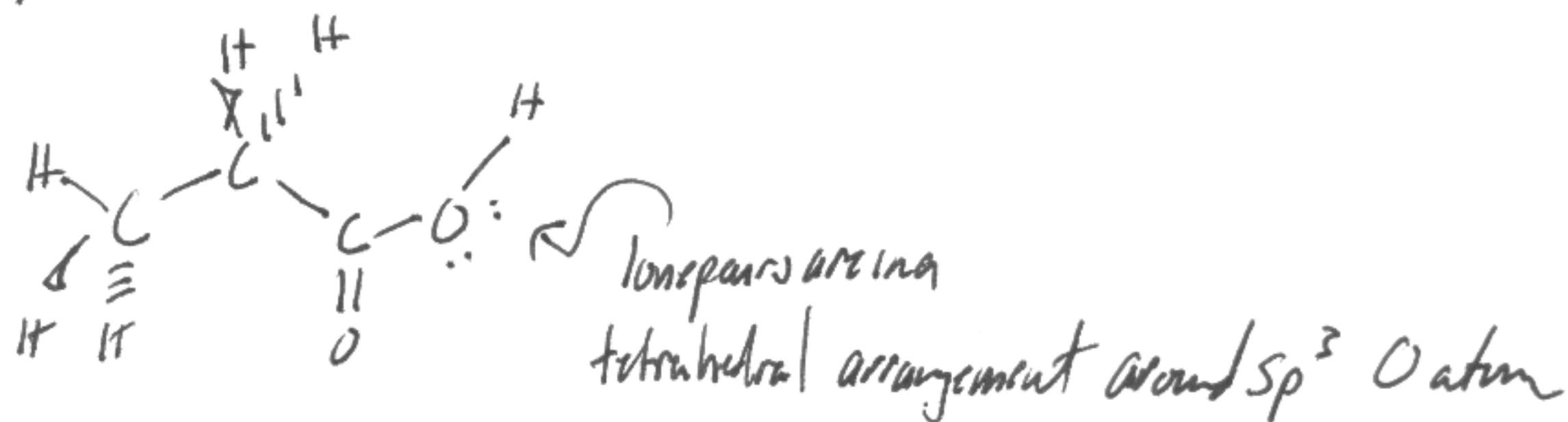
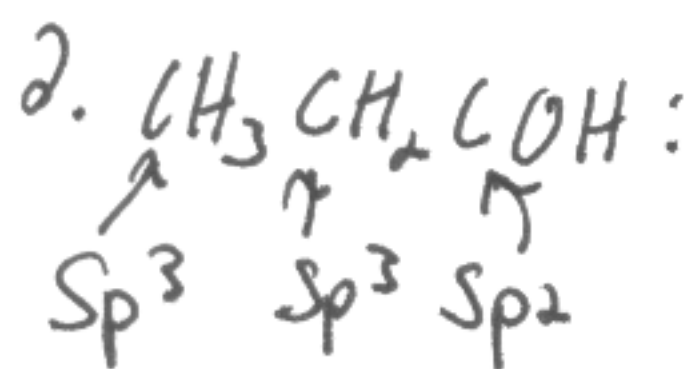
Homework week 12, key



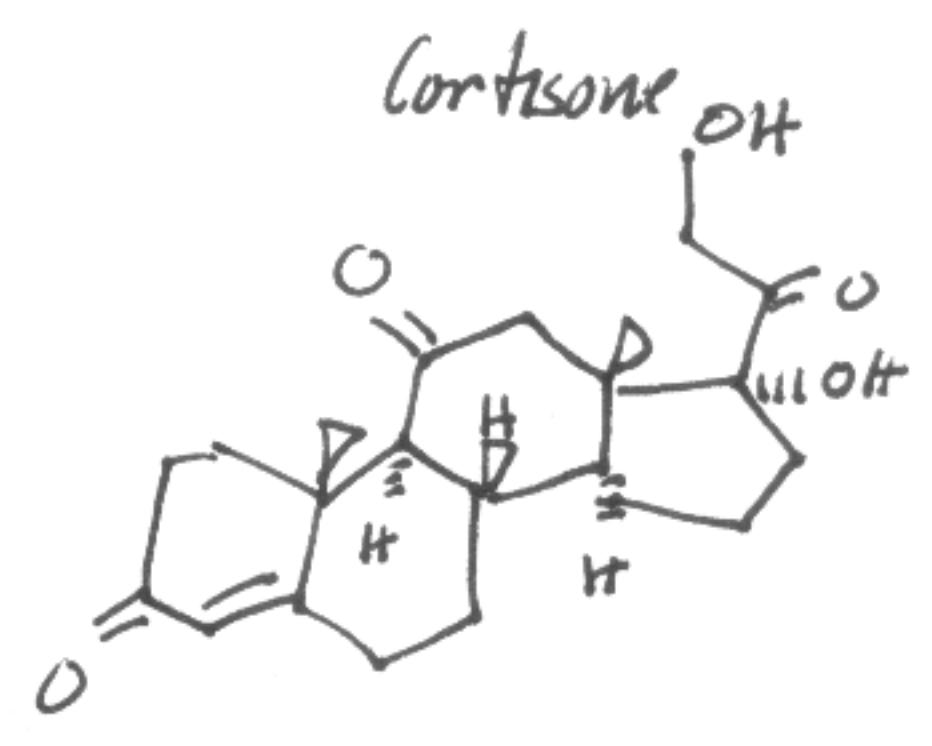
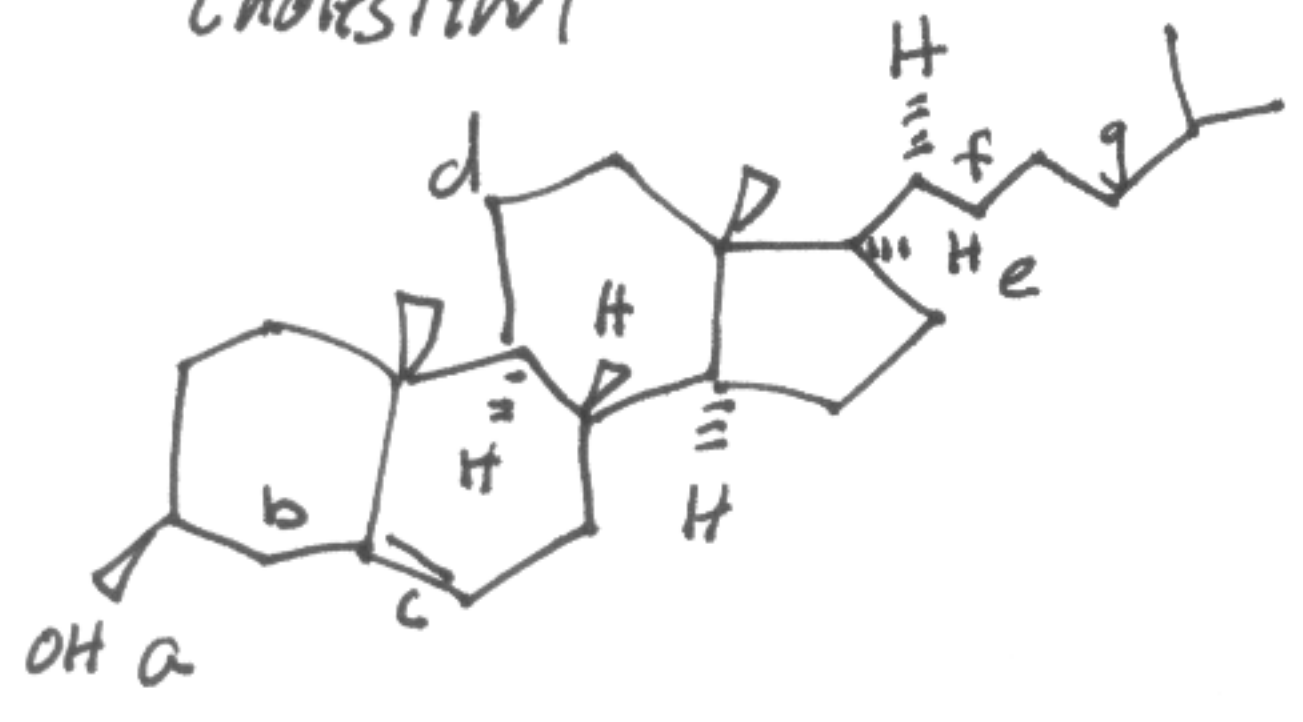
In both cases the C atom is sp^2 hybridized. In the acid, 1 O atom is sp^2 and the other is sp^3 . However, in the ion, resonance structures flip the sp^2-sp^3 hybridization of the 2 O atoms. This means that both resonance structures will have some character of sp^2

hybridization, and that the  functional group will remain trigonal planar. This also means that the bond lengths of the C-O bonds will be intermediate between the two extremes of C-O and C=O.

Assume the average, 1.8 \AA .

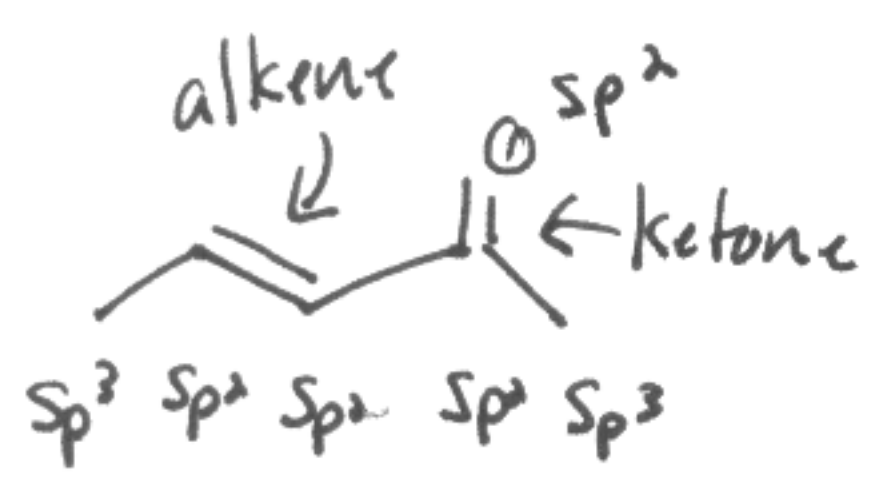


3. Cholesterol

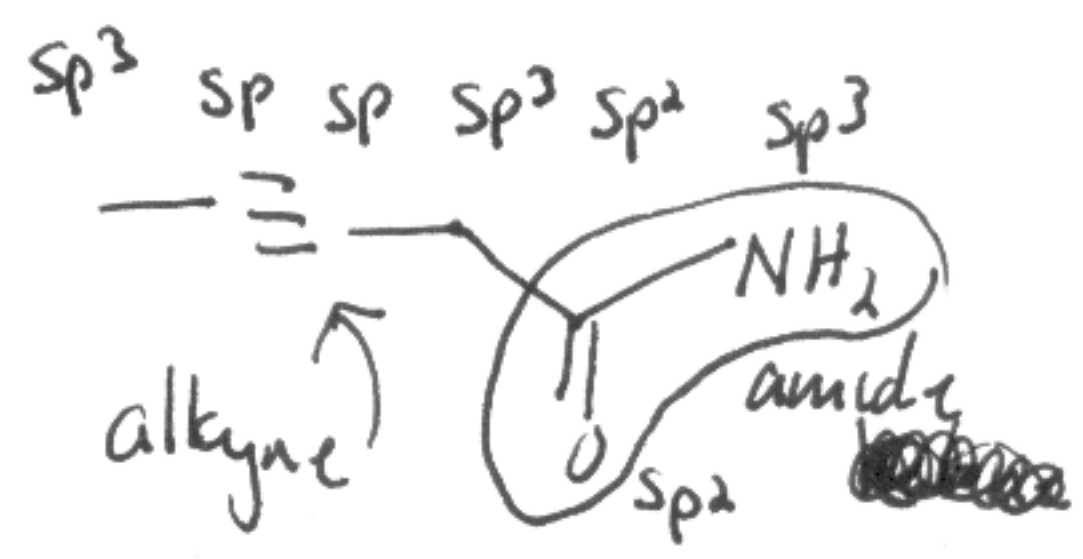


- a) convert alcohol to ketone
- b) convert alkane to alkene
- c) convert alkene to alkane
- d) convert alkane to ketone
- e) convert alkane to alcohol
- f) convert alkane to ketone
- g) convert alkane to alcohol

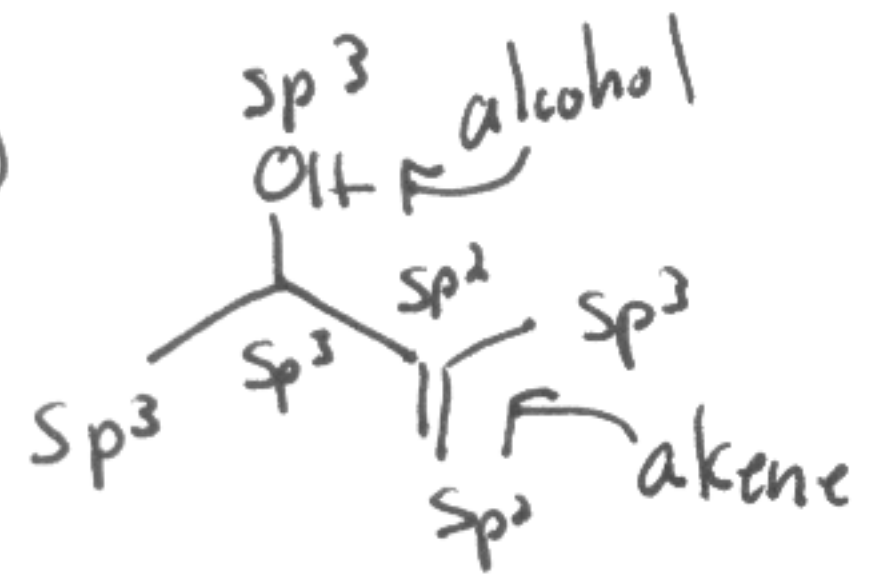
4. a)



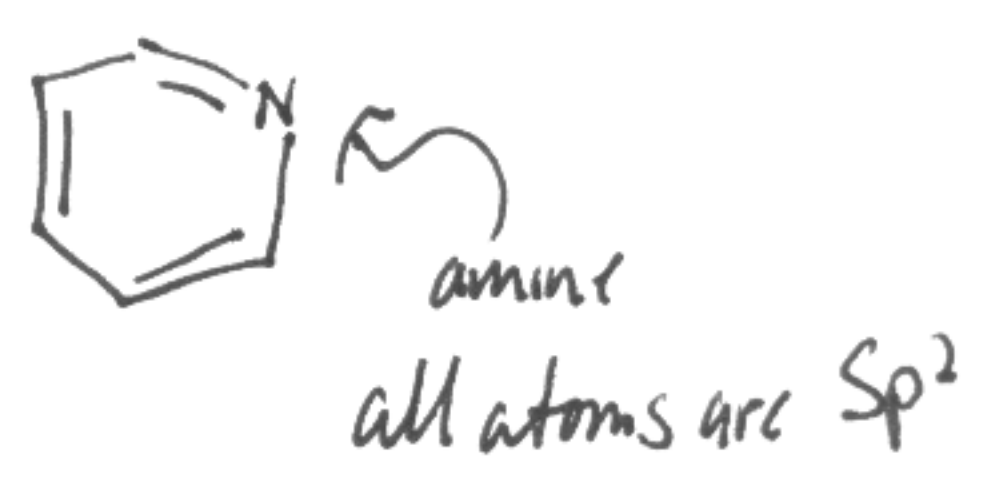
b)

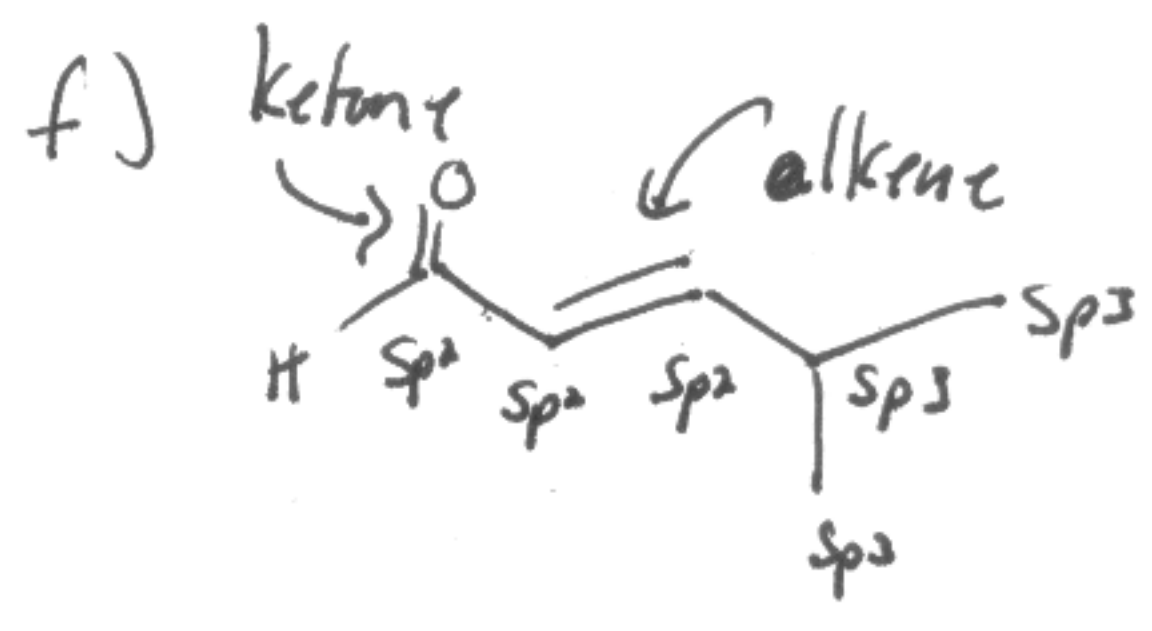
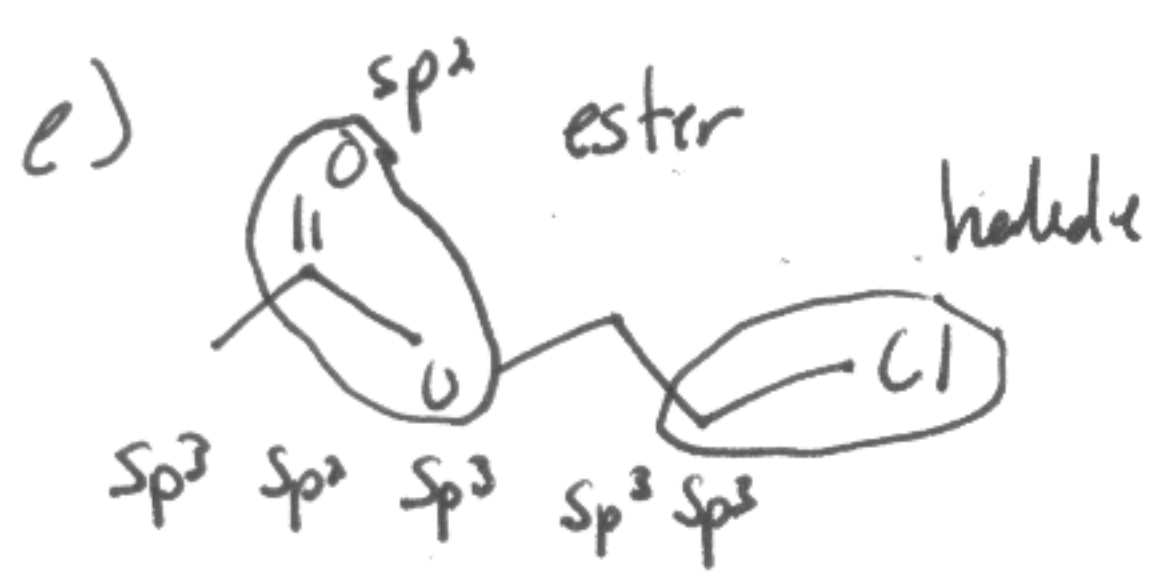


c)



d)





S. $m = 250g$ of CO_2

$FW(CO_2) = 44.01 g/mol$

$V = 1.5 dm^3 = 1.5 L$

$T = 25^\circ C = 298 K$

$n_{CO_2} = \frac{250g}{44.01g/mol} = 5.68 mol$

if ideal: $PV = nRT$

$P = \frac{nRT}{V} = \frac{(5.68 mol)(8.20 \times 10^{-2} Latm/mol K)(298 K)}{1.5 L}$

$P = 92.7 atm$

if vdW: $P = \frac{nRT}{V-nb} - \frac{an^2}{V^2}$

$a = 3.6073 L^2 atm/mol^2$

$b = 0.0428 L/mol$

$P = \frac{(5.68 mol)(8.20 \times 10^{-2} Latm/mol K)(298 K)}{1.5 L - (5.68 mol)(0.0428 L/mol)} - \frac{(3.6073 L^2 atm/mol^2)(5.68 mol)^2}{(1.5 L)^2}$

$P = 58.8 atm$

Under these conditions, attractive forces dominate the gas.

(4)

$$6. V_L = 2.0L$$

$$P_L = 3.0 \text{ atm}$$

$$V_f = 2.0L + 5.0L = 7.0L$$

$$P_L V_L = P_f V_f ; \quad P_f = \frac{P_L V_L}{V_f} = \frac{(3.0 \text{ atm})(2.0L)}{7.0L}$$

$$P_f = 0.86 \text{ atm}$$

$$7. \text{ He: } c_{rms} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3(8.31 \text{ J/kmol})(6000 \text{ K})}{0.004003 \text{ kg/mol}}} = 6100 \text{ m/s} \quad (6000 \text{ K})$$

$$\text{@ } 100 \text{ K: } c_{rms} = \sqrt{\frac{3(8.31 \text{ J/kmol})(100 \text{ K})}{0.004003 \text{ kg/mol}}} = 790 \text{ m/s}$$

SINCE c_{rms} INCREASES AS $\propto \sqrt{T}$, EVEN THOUGH T HAS GONE UP BY A FACTOR OF 60, c_{rms} HAS INCREASED BY A FACTOR OF ≈ 8 .